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Consider a 0.238 M aqueous solution of sodium hydroxide, NaOH. a) How many grams of NaOH are dissolved in 23.46 mL? b) How many individual hydroxide ions ( $\text{OH}^-$ ) are found in 23.46 mL? c) How many moles of sulfuric acid,  $\text{H}_2\text{SO}_4$ , are neutralized by 23.46 mL of 0.238 M NaOH(aq)? [Hint: begin by writing a balanced equation for this neutralization reaction.]

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\_\_\_ Fe+2 → \_\_\_ Mn+2 \_\_\_ Fe+3 \_\_\_ mLs  
20 mLs 0.01 M ?

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